## Section-B

## (Short Answer)

**Note: Answer any EIGHT of the following questions. Each question carries 05 marks.**

Q.2 Define chemistry and describe its importance.

Q.3 What is the difference between Empirical formula and Molecular formula?

Q.4 State five uses of Hydrogen.

Q.5 What do you understand by long form of periodic table?

Q.6 Write the characteristics of covalent compound.

Q.7 What do you mean by liquid? Describe the interchange of liquid to gas.

Q.8 Differentiate between Solution and Suspension.

Q.9 State the structure of Diamond or Graphite.

Q.10 Define heat of neutralization. What would be the value of heat of neutralization, when strong acid reacts with strong base?

Q.11 Nitrogen is one of the major components of air. How can you get nitrogn from the atomospheric air?

Q.12 What is plastic sulphur and why is it elastic?

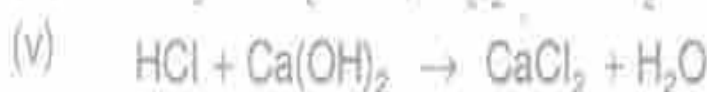
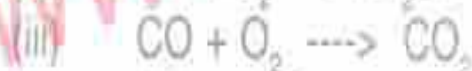
Q.13 Balance the folloiwng equation:

(i) $NaOH + H_2SO_4 \rightarrow NaSO_4 + H_2O$

(ii) $Na + O_2$ ---> $Na_2O$

(iii) $CO + O_2$ ----> $CO_2$

(iv) $NH_3 + CO_2$ ---> $(NH_2)_2\,CO + H_2O$

(v) $HCl + Ca(OH)_2 \rightarrow CaCl_2 + H_2O$

## Section-C

## (Descriptive Answer)

**Note: Answer any TWO of the following question.**

Q.14(a) Descriobe the construction and working of dry cell.

(b) The pOH of a solution is 9.40 Calculate the $(H^+)$ ion concentration.

Q.15(a): State the Law of Reciprocal Proportion and illustrate with example.

(b) What are distinguishing characters of electron , proton and neutron?

Q.16(a): How is chlorine prepared in the laboratory?

(b) Explain the main sources of organic compounds.